

Warm Up:

What are the names of the following compounds?



Warm Up

Write the formulas for the following:

Strontium cyanide

Mercury (I) phosphate

Copper (II) sulfite

Objectives:

TSWBAT:

- **Name binary chemical compounds.**
- **Name compounds containing polyatomic ions**

10 Polyatomic Ions for the Quiz

1. Nitrite
2. Nitrate
3. Cyanide
4. Hydroxide
5. Sulfite
6. Sulfate
7. Carbonate
8. Phosphite
9. Phosphate
10. Ammonium

Aluminum iodide



Calcium sulfide



Lithium phosphide



Strontium bromide



Aluminum selenide



Potassium nitride



Beryllium chloride



Sodium oxide



Aluminum fluoride



Cesium sulfide



Barium phosphide



Francium oxide



Radium nitride



Magnesium bromide



Main Block
Ionic Cmpds

Transition metal compounds

Iron (II) iodide FeI_2

Copper (I) sulfide Cu_2S

Tin (II) phosphide Sn_3P_2

Lead (IV) bromide PbBr_4

Mercury (II) selenide HgSe

Copper (II) nitride Cu_3N_2

Iron (III) chloride FeCl_3

Lead (IV) oxide PbO_2

Tin (II) fluoride SnF_2

Iron (III) sulfide Fe_2S_3

Mercury (I) phosphide Hg_3P

Copper (II) oxide CuO

Mercury (II) nitride Hg_3N_2

Copper (I) bromide CuBr

Polyatomic Compounds

Beryllium nitrate $\text{Be}(\text{NO}_3)_2$

Iron (III) carbonate $\text{Fe}_2(\text{CO}_3)_3$

Sodium phosphate Na_3PO_4

Calcium chlorate $\text{Ca}(\text{ClO}_3)_2$

Copper (II) hydroxide $\text{Cu}(\text{OH})_2$

Ammonium phosphate $(\text{NH}_4)_3\text{PO}_4$

Mercury (I) sulfate Hg_2SO_4

Aluminum hydroxide $\text{Al}(\text{OH})_3$

Magnesium nitrate $\text{Mg}(\text{NO}_3)_2$

Tin (II) carbonate SnCO_3

Strontium chlorate $\text{Sr}(\text{ClO}_3)_2$

Lithium phosphate Li_3PO_4

Lead (II) sulfate PbSO_4

Ammonium sulfide $(\text{NH}_4)_2\text{S}$

Metal Compounds

Lead (IV) oxide

Tin (II) fluoride

Iron (III) sulfide

Mercury (I) phosphide

Copper (II) oxide

Mercury (II) nitride

Copper (I) bromide

Practice Sheet

Polyatomic Ions

Calcium nitrite

Magnesium nitrate

Lithium Cyanide

Potassium hydroxide

Rubidium sulfite

Beryllium sulfate

Strontium carbonate

Barium phosphite

Magnesium phosphate

Cesium Nitrate

Francium Nitrate

Polyatomic ions, continued:

sodium cyanide

aluminum hydroxide

barium sulfite

gallium sulfate

beryllium carbonate

magnesium phosphite

rubidium phosphate

strontium nitrate

lithium phosphate

sodium sulfide

aluminum sulfite

Practice Sheet Stock System

Copper (I) oxide

Mercury (I) bromide

Nickel (II) selenide

Lead (IV) fluoride

Cadmium oxide

Gold (III) phosphide

Nickel (III) iodide

Zinc Nitride

Silver chloride

Cobalt (III) sulfide

Tin (II) bromide

Copper (II) oxide

Lead (II) oxide

Stock System Practice, continued

Chromium (III) selenide

Iron (II) phosphide

Gold (I) iodide

**Name the following
chemical compounds**



Did you do your homework? (Front of sheet?)

Warm Up

Write formulas for these compounds.

- a. sodium iodide b. Tin (II) chloride
c. potassium sulfide d. calcium iodide

Objectives:

TSWBAT:

- **Name binary chemical compounds.**
- **Name compounds containing polyatomic ions**

Naming Binary Compounds

Binary compounds are compounds composed of **two** elements. They may be ionic or covalent depending on which elements are combined.